

Centre Number	Candidate Number	Name
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UNIVERSITY OF CAMBRIDGE INTERNATIONAL EXAMINATIONS
International General Certificate of Secondary Education

PHYSICAL SCIENCE

0652/03

Paper 3

May/June 2004

1 hour 15 minutes

Candidates answer on the Question Paper.
No Additional Materials are required.

READ THESE INSTRUCTIONS FIRST

Write your Centre number, candidate number and name on all the work you hand in.
Write in dark blue or black pen in the spaces provided on the Question Paper.
You may use a pencil for any diagrams, graphs, tables or rough working.
Do not use staples, paper clips, highlighters, glue or correction fluid.

Answer **all** questions.

At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

A copy of the Periodic Table is printed on page 16.

For Examiner's Use

1

2

3

4

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6

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9

Total

If you have been given a label, look at the details. If any details are incorrect or missing, please fill in your correct details in the space given at the top of this page.

Stick your personal label here, if provided.

Answer **all** the questions.

Write your answers in the spaces provided.

- 1 The metal **M** has a relative atomic mass, A_r , of 30.
M is reacted with pure oxygen at room temperature and pressure (r.t.p.) to form the oxide of **M**.
 It is found that 5.0 g of **M** react with exactly 2.0 dm³ of oxygen under these conditions.
 1 mole of any gas occupies 24 dm³ at r.t.p.

(a) Define the term *relative atomic mass*.

.....

[2]

(b) (i) Find the number of moles in 5.0 g of **M**.

number of moles =[2]

(ii) Find the number of moles of oxygen gas that react with 5.0 g of **M**.

number of moles =[1]

(iii) From your answers to (i) and (ii) deduce the number of moles of **M** that react with 1 mole of oxygen gas.

number of moles =[1]

(iv) Write the balanced chemical equation for the reaction between **M** and oxygen.

.....[2]

(b) Fig. 2.2 is a table of the density of some metals.

aluminium	2.7
copper	8.9
gold	19.3
iron	7.8
lead	11.3
magnesium	1.7
platinum	21.4
uranium	18.7

Fig. 2.2

(i) State a unit for density.

.....[1]

(ii) A sample of metal of measured volume 0.8 cm^3 is found to have a mass of 15.4 g . Calculate the density of the sample. Write down the equation that you use and show your working.

[2]

(iii) Which of the metals in the table is the sample most likely to be?

.....[1]

(iv) Explain why the value calculated in (ii) may not be exactly the same as any of the values in the table.

.....

[2]

- (c) A second sample of the metal is found to have a mass of 85 g. On the moon it is found to weigh 0.14 N. Calculate the value of the gravitational field strength (g) on the moon. Write down the equation that you use and show your working.

value for g on the moon = [3]

- 3 Fig. 3.1 contains the melting points of the elements in Period 3 of the Periodic Table.

element	Na	Mg	Al	Si	P	S	Cl	Ar
m.pt./K	371	924	933	1680	317	386	172	84

Fig. 3.1

- (a) Describe the trend in melting point of the elements in Period 3.

.....
[1]

- (b) What does the high melting point of silicon suggest about the structure of solid silicon?

.....

[2]

- (c) Use your knowledge of the trends in the Periodic Table to deduce which of these elements

- (i) is the most reactive metal,[1]
 (ii) has five electrons in the outer shell of one atom,[1]
 (iii) forms an ion with a charge of 2+,[1]
 (iv) does not react with any other element.[1]

- (d) Suggest why sodium is a softer metal than magnesium.

.....

.....

.....

.....[2]

- 4 Fig. 4.1 shows a circuit used to measure the resistance of different lengths of wire.

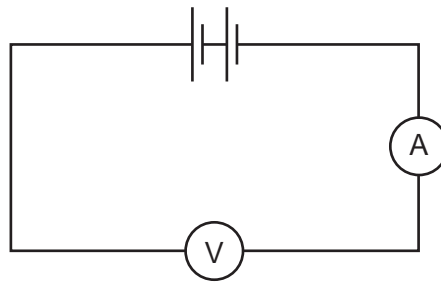


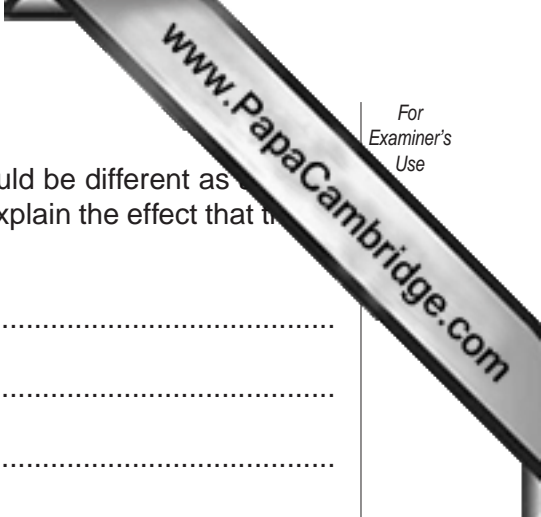
Fig. 4.1

- (a) Mark on Fig. 4.1 where the resistance wire should be connected. [1]
- (b) The voltmeter shows a value of 4.3 V and the ammeter a value of 2.1 A.
Calculate the resistance of the wire.
Write down the equation that you use and show your working.

resistance of wire =[3]

- (c) What would you expect the resistance to be when a wire of the same metal and same diameter but of twice the length replaces the one used in (b)?

new resistance =[1]



(d) It is likely that the temperatures of the two lengths of wire would be different as a result of the different currents flowing through the wires. State and explain the effect that this is likely to have on the resistances of the wires.

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.....
.....
.....
.....
.....
.....[3]

(e) Explain why using a very short length of resistance wire could damage the ammeter.

.....
.....
.....
.....[2]

(f) Name a device that could be used to measure the p.d. instead of the voltmeter.

.....[1]

5 (a) (i) Using the Periodic Table on page 16, write down the electronic structures of calcium and fluorine.

electronic structure of calcium[1]

electronic structure of fluorine[1]

(ii) By reference to these structures, describe how calcium and fluorine form calcium fluoride.

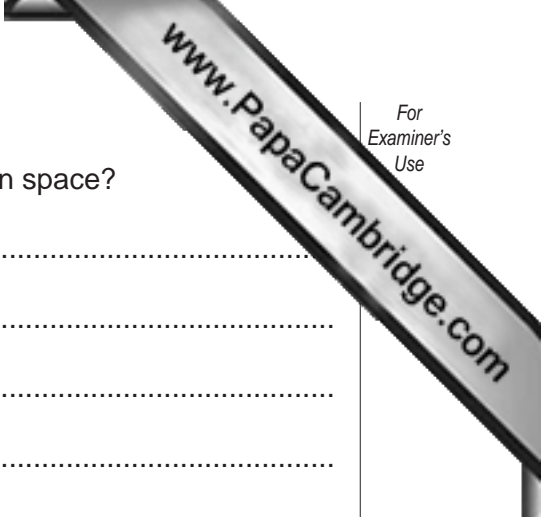
[2]

(iii) State the formula of calcium fluoride.

.....[1]

(b) Solid calcium fluoride does not conduct electricity but molten calcium fluoride does. Liquid fluorine does not conduct electricity. Explain these observations in terms of the structures of these substances.

.....
.....
.....
.....
.....
.....[3]



(ii) Why is it important to carry out this experiment in an open space?

.....
.....
.....
.....
.....
.....

[2]

(d) A particular sound wave has a frequency of 3.5 kHz. What is its wavelength?
Write down the equation that you use and show your working.

wavelength of sound =[3]

- 7 When an aqueous solution of sugars is fermented to form ethanol the resulting solution contains about 15% of ethanol.
- (a) State two essential conditions for fermentation to occur.
1.
2.[2]
- (b) (i) The boiling point of ethanol is 78 °C.
State the name of the method used to separate ethanol from the solution.
.....[1]
- (ii) Sketch a labelled diagram of the apparatus that would be used in the laboratory to carry out this separation.

[4]

8 (a) Complete the following passage relating to measuring temperatures.

Temperatures indicate the degree of "hotness" of an object and can be measured using a These are devices that must have a property which with temperature. Ideally this property shows a linear relationship with temperature. This means that the property changes by equal amounts for changes in temperature. The maximum temperature minus the minimum temperature that the device can register is known as the of the device. Devices for which the chosen property makes a large change for a small change in temperature are said to be very [5]

(b) Name a specific type of temperature measuring device and the property on which its operation depends.

device
 property [2]

9 Three substances are added to a blast furnace during the extraction of iron from its ore. Haematite, Fe_2O_3 , is the ore. Limestone, $CaCO_3$, and coke are the other substances needed. The main products are iron, slag and oxides of carbon.

(a) State the reason for adding limestone to the furnace.
 [1]

(b) Write balanced equations for the reactions that occur when
 (i) limestone decomposes,
 (ii) haematite is reduced. [4]

(c) Iron and steel can be prevented from rusting by galvanising. When the covering is complete this prevents the iron being exposed to the air. However, unlike painting, galvanising continues to prevent rusting even when the coating is damaged. Explain this difference between painting and galvanising.

 [2]

15
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DATA SHEET
The Periodic Table of the Elements

Group

I	II	III	IV	V	VI	VII	O		
7 Li Lithium 3	9 Be Beryllium 4	1 H Hydrogen 1	11 B Boron 5	12 C Carbon 6	14 N Nitrogen 7	16 O Oxygen 8	19 F Fluorine 9	20 Ne Neon 10	
23 Na Sodium 11	24 Mg Magnesium 12	27 Al Aluminium 13	28 Si Silicon 14	31 P Phosphorus 15	32 S Sulphur 16	35.5 Cl Chlorine 17	40 Ar Argon 18		
39 K Potassium 19	40 Ca Calcium 20	45 Sc Scandium 21	51 V Vanadium 23	59 Co Cobalt 27	65 Zn Zinc 30	70 Ga Gallium 31	79 Se Selenium 34	84 Kr Krypton 36	
85 Rb Rubidium 37	88 Sr Strontium 38	89 Y Yttrium 39	93 Nb Niobium 41	103 Rh Rhodium 45	112 Cd Cadmium 48	115 In Indium 49	128 Te Tellurium 52	131 Xe Xenon 54	
133 Cs Caesium 55	137 Ba Barium 56	139 La Lanthanum 57	181 Ta Tantalum 73	192 Ir Iridium 77	201 Hg Mercury 80	204 Tl Thallium 81	207 Pb Lead 82	209 Bi Bismuth 83	210 Po Polonium 84
226 Ra Radium 88	227 Ac Actinium 89	227 Fr Francium 87	227 Fr Francium 87	227 Ac Actinium 89	227 Fr Francium 87	227 Ac Actinium 89	227 Fr Francium 87	227 Ac Actinium 89	227 Fr Francium 87

140 Ce Cerium 58	141 Pr Praseodymium 59	144 Nd Neodymium 60	150 Sm Samarium 62	152 Eu Europium 63	157 Gd Gadolinium 64	162 Dy Dysprosium 66	165 Ho Holmium 67	167 Er Erbium 68	169 Tm Thulium 69	173 Yb Ytterbium 70	175 Lu Lutetium 71	
232 Th Thorium 90	232 Pa Protactinium 91	238 U Uranium 92	238 Pu Plutonium 94	238 Np Neptunium 93	238 Am Americium 95	238 Cm Curium 96	238 Bk Berkelium 97	238 Cf Californium 98	238 Es Einsteinium 99	238 Fm Fermium 100	238 Md Mendelevium 101	238 No Nobelium 102

8-71 Lanthanoid series
10-103 Actinoid series

a = relative atomic mass
X = atomic symbol
b = proton (atomic) number

The volume of one mole of any gas is 24 dm³ at room temperature and pressure (r.t.p.).